

In summary, we have demonstrated that carbon monoxide can be efficiently trapped by carbon radicals at reasonably low CO pressures, contrary to previous observations. The success of this *free-radical carbonylation* obviously provides a new method for the introduction of carbon monoxide into organic molecules.^{9,10}

(9) For the preexisting methodologies for carbonylation with CO involving other acyl species (acylmethyls, acyl anions, and acyl cations), see recent reviews: (a) Narayama, C.; Periasamy, M. *Synthesis* **1985**, 253. (b) Weil, T. A.; Casser, L.; Foa, M. In *Organic Synthesis via Metal Carbonyls*; Wender, I., Pino, P., Eds.; Wiley: New York, 1977; Vol. 2, p 517. (c) Bahrman, H.; Cornils, B.; Frohling, C. D.; Mullen, A. In *New Syntheses with Carbon Monoxide*; Falbe, J., Ed.; Springer: Berlin, 1980.

(10) The overall transformation as a method for aldehyde synthesis is noteworthy, since the carbonylation of aliphatic substrates is particularly difficult by transition-metal methods. In this context, the radical method reported here complements the method involving Pd catalysis by J. K. Stille; see: Baillargeon, V. P.; Stille, J. K. *J. Am. Chem. Soc.* **1983**, *105*, 7175.

Further studies on the scope and synthetic application are currently under investigation.

Registry No. **1**, 111-83-1; **2**, 124-19-6; **3**, 111-65-9; Ph(CH₂)₃Br, 637-59-2; Ph(CH₂)₂Br, 103-63-9; PhCH₂Br, 100-39-0; PhI, 591-50-4; *p*-MeOC₆H₄I, 696-62-8; *c*-C₆H₁₁Br, 108-85-0; (*E*)-EtCH=CH(CH₂)₂Br, 63281-96-9; (*Z*)-EtCH=CH(CH₂)₂Br, 5009-31-4; Ph(CH₂)₃CHO, 18328-11-5; Ph(CH₂)₂CHO, 104-53-0; PhCH₂CHO, 108-88-3; PhCHO, 100-52-7; *p*-MeOC₆H₄CHO, 123-11-5; *c*-C₆H₁₁CHO, 2043-61-0; (*E*)-EtCH=CH(CH₂)₂CHO, 929-22-6; (*Z*)-EtCH=CH(CH₂)₂CHO, 6728-31-0; CO, 630-08-0; 1-adamantyl bromide, 768-90-1; 1-adamantanecarboxaldehyde, 2094-74-8.

Supplementary Material Available: ¹H and ¹³C NMR, IR, and some mass spectral data for products listed in Table I (3 pages). Ordering information is given on any current masthead page.

Additions and Corrections

Thermal Encapsulation and Photochemical Deencapsulation of Ag(I) by [Ir₂(dimen)₄](PF₆)₂ (dimen = 1,8-Diisocyanomethane). X-ray Crystal Structure of [AgIr₂(dimen)₄](PF₆)₃·2DMSO [J. Am. Chem. Soc. **1988, *110*, 8252]. ANDREW SYKES and KENT R. MANN***

Page 8253: In Figure 2, the formation constant for the thermal encapsulation of Ag⁺ in DMSO is misstated as 1.5 × 10⁸ M⁻¹. The correct value is 1.5 × 10⁷ M⁻¹. The calculations in the text (page 8253) should be adjusted to give log *K* of 7.2 and the corresponding free energy change of -9.8 kcal/mol.

Time-Resolved Raman Detection of ν(Fe-O) in an Early Intermediate in the Reduction of O₂ by Cytochrome Oxidase [J. Am. Chem. Soc. **1989, *111*, 6439-6440]. CONSTANTINOS VAROTSIS, WILLIAM H. WOODRUFF, and GERALD T. BABCOCK***

The spectrometer used in this characterization of the dioxygen adduct of cytochrome *a*₃²⁺ in the reaction of fully reduced cytochrome oxidase with O₂ was miscalibrated in the 540-660 cm⁻¹ region. With proper calibration, the 589- and 565-cm⁻¹ lines occur at 571 and 546 cm⁻¹, respectively. Therefore, we assign the ν(Fe²⁺-O) frequency at 571 cm⁻¹. This value is very similar to that which we measured for the iron-oxygen stretching frequency in an imidazole-heme *a*²⁺-O₂ model compound (see ref 7 in the original publication) and identical with that which we measured in the reaction of mixed valence cytochrome oxidase with O₂. (Varotsis, C.; Woodruff, W. H.; Babcock, G. T. *J. Biological Chem.* Submitted).

This indicates, contrary to our earlier conclusion, that the cytochrome *a*₃-O₂ complex is unperturbed by distal effects in the cytochrome *a*₃/Cu_B binding pocket. Weakening and rupture of the O=O bond occurs subsequent to formation of the initial dioxygen-*a*₃²⁺ adduct.

Selenium Coronands: Synthesis and Conformational Analysis [J. Am. Chem. Soc. **1989, *111*, 6582]. RAYMOND J. BATCHELOR, FREDERICK W. B. EINSTEIN, IAN D. GAY, JIAN-HUA GU, BLAIR D. JOHNSTON, and B. MARIO PINTO***

Recent investigations in our laboratory show that the solid-state ⁷⁷Se chemical shifts reported in ref 1 are incorrect. This arises from an error in the referencing procedure. Our measurements were referred to a solution of aqueous H₂SeO₃ and converted to the (CH₃)₂Se scale with use of the literature value^{2,3} of 1282 ppm

for the shift of H₂SeO₃. Direct measurement shows this figure to be incorrect, and the shifts reported in Table XI of ref 1 should be corrected by +22 ppm to give shifts relative to (CH₃)₂Se in CDCl₃.

Page 6584: In footnotes *i* and *j* to Table II *W* should be *w*.

Page 6588 first column, line 25: "clockwise" should be "counterclockwise".

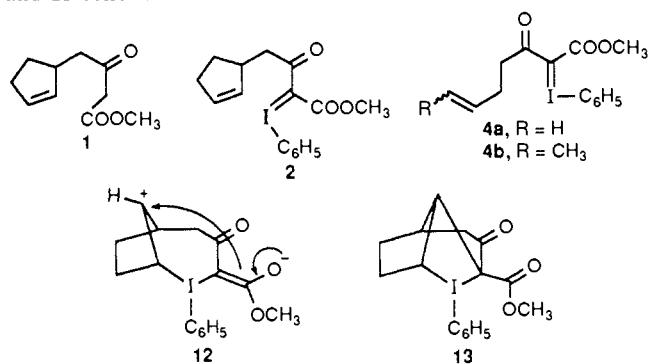
(1) Batchelor, R. J.; Einstein, F. W. B.; Gay, I. D.; Gu, J.-H.; Johnston, B. D.; Pinto, B. M. *J. Am. Chem. Soc.* **1989**, *111*, 6582.

(2) *NMR and the Periodic Table*; Harris, R. K., Mann, B. E., Eds.; Academic Press: New York, 1978.

(3) Lardon, M. *J. Am. Chem. Soc.* **1970**, *92*, 5063.

A Novel Intramolecular Cyclopropanation Using Iodonium Ylides [J. Am. Chem. Soc. **1989, *111*, 6443]. ROBERT M. MORIARTY,* OM PRAKASH, RADHE K. VAID, and LEI ZHAO**

Pages 6443 and 6444: The correct structures for **1**, **2**, **4**, **12**, and **13** follow:



Page 6443: The following should be added to ref 10. **4a:** ¹H NMR (CDCl₃) δ 2.50 (t, 2 H, CH₂), 3.20 (m, 2 H, CH₂), 3.75 (s, 3 H, COOCH₃), 5.10 (m, 2 H, CH=CH₂), 6.09 (m, 1 H, CH=CH₂), 7.40-7.90 (m, 5 H, aromatic protons). **6:** ¹H NMR (CDCl₃) δ 1.30-2.10 (m, 4 H, 2 × CH₂), 2.85 (m, 1 H, CH), 3.20 (m, 1 H, CH), 3.70 (s, 3 H, COOCH₃), 4.15 (m, 1 H, CH), 5.80 (m, 2 H, CH=CH), 6.20 (m, 1 H, CH=CH), 7.30-7.90 (m, 5 H, aromatic protons).